1. MSDS, Safety, Etc.
   b. Know MSDS Information for the chemicals used in the experiments below (to include):
      - **Heat of Neutralization**: HCl, NaOH, ice
      - **Radiochemistry**: α-, β-, and γ- radiation
      - **Antacid Analysis**: Phenolphthalein, CaCO₃, Al(OH)₃, MgCO₃, HCl & NaOH
      - **Spectrophotometry & Colorimetry**: water ☺
      - **Equilibrium/Beer’s Law**: Fe(NO₃)₃, HNO₃, KSCN
      - **Atomic Spectra**: FeCl₂, LiCl, KCl, NaCl, CuCl₂, SrCl₂, CH₂Cl₂, Fe(NO₃)₂, LiNO₃, KNO₃, NaNO₃, Cu(NO₃)₂
      - **Gas Laws**: Acetic acid (CH₃COOH), Baking soda (NaHCO₃), Butane
   c. **Room Diagram** – Be able to label equipment in your room.
      - Balances, Distilled Water, Exits, Fire Extinguishers, Hoods,
      - Safety Blankets, Safety Showers, Waste Containers

2. **Heat of Neutralization / Heat of Fusion**
   a. Read over the green book pages 1-18. For more information see Averill Chem 1 text chapter 5.2-3 pp. 209-229.
   b. Be able to balance an acid base reaction equation.
   c. Be able to determine the limiting reagent, given the molarity of the acid/base and
      the number of ml of each.
   d. Be able to determine the enthalpy of a reaction given the heat capacity.
      \[ \Delta H = \frac{-\text{total } C_p \times \Delta T}{\text{moles reacted}} \]

3. **Radiochemistry**
   b. Be able to balance nuclear decay equations for α-emission, β-emission, and neutron emission.
   c. If given the time and counts, be able to find the natural log of the counts (ln counts).
   d. Be able to determine the specific decay constant, k, by finding the slope of a line.
      (Note: it is always best to use data points furthest apart to determine the slope of a line.)
      The slope of a line:
      \[ m = \frac{(y_2 - y_1)}{(x_2 - x_1)} \]
   e. Having calculated the slope and using any data point, be able to find the y intercept, b, of a line:
      \[ y = mx + b \]
      \[ b = y - mx \]
   f. Having found the y-intercept, b, be able to convert the answer from ln counts to counts in
      order to find A₀.
   g. Having calculated k, be able to determine the half-life of the compound.
   h. Be able to calculate the percent error (percent difference) of the calculated half-life vs. a
      given theoretical half-life.

4. **Antacid Analysis**
   a. Read over the green book pp 33-44.
   b. Know how to balance equations for antacids reacting with HCl.
   c. Given concentrations and volumes of HCl and NaOH, know how to determine how much acid was
      neutralized by the antacid.
   d. Be able to determine how much acid the antacid should have been able to theoretically neutralize.
5. Spectrophotometry and Colorimetry
   b. Know the equation for Absorbance and how to convert from transmittance to absorbance.
      \[ A = \log \left( \frac{100}{\%T} \right) \]
   c. Know how to find the maximum absorbance for individual unknowns and for a mixture of colors.
   d. Be able to calculate the absorbance ratio of an unknown vs. a standard.
   e. Be able to calculate the concentration of the unknowns if given the concentration of the standards.

6. Equilibrium/Beer’s Law
   a. Read over Handout.
   b. Know how to determine \( K_{eq} \) from a graph. \( K_{eq} = m \) for \( y = mx + b \), when
      \[ x = \frac{A(\text{[Fe}^\ast\text{]} + \text{[SCN}^\ast\text{]})}{(\text{[Fe}^\ast\text{]} \text{[SCN}^\ast\text{]})} \]
      \[ y = \frac{A}{\text{[Fe}^\ast\text{]} \text{[SCN}^\ast\text{]}} \]
   c. Know how to calculate the concentration of an unknown solution given the %T.
      \[ A = \log \left( \frac{100}{\%T} \right) \]
      \[ A = abc \]

7. Atomic Spectra
   b. Know the Rydberg equation and thus how to calculate frequency, \( \nu \).
      \[ \nu = R \left( \frac{1}{n_1^2} - \frac{1}{n_2^2} \right) \]
   c. Know how to convert from frequency, \( \nu \), to wavelength, \( \lambda \), using the speed of light, \( C \).
      \[ C = \lambda \nu \]
   d. Know which wavelengths correspond to the Balmer series (visible) and which ones correspond to the Lyman series (ultraviolet).

8. Gas Laws and Buoyancy Effects
   b. Know the equation for the Ideal Gas Law:
      \[ PV = nRT \]
   c. Realize that given the gas law constants that individual data must correspond for units to cancel. For example, if \( R \) is in units atm \( \cdot \) L / mole \( \cdot \) K then \( T \) must be in K not \( ^\circ \)C.
      \[ \text{Pressure: } 760 \text{ torr} = 1 \text{ atm} \]
      \[ \text{Volume: } 1000 \text{ml} = 1 \text{ L} \]
      \[ \text{Temperature: } ^\circ \text{C} + 273.15 = K \]
      \[ \text{Moles: } MW = g / \text{mole} \]
   d. Know how to set up an equilibrium in order to convert from one set of conditions to another, if one condition (pressure, volume, temperature or number of moles) is altered.
      For the Ideal Gas Law
      \[ \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2} \]
   e. Know the equation for density (\( d = m / v \)) and how to use it to convert from mass to volume or volume to mass.

9. Statistical Analysis
   b. Know how to calculate the mean (average) of a set of data.
      1. Average or mean:
      \[ \bar{x} = \frac{\sum x_i}{n} \]
   c. Know how to calculate the standard deviation or estimate, if given the equation:
      2. Standard Deviation:
      \[ \sigma = \sqrt{\frac{\sum (x_i - \bar{x})^2}{n}} \]
      3. Estimate of the Standard Deviation:
      \[ s = \sqrt{\frac{\sum (x_i - \bar{x})^2}{n-1}} \]
   d. Know the differences between equations 2 & 3 and when each of these equations is applicable.

10. Dimensional Analysis & Scientific Notation
    b. Be able to evaluate problems similar to those in each of the following sections.
       1. Problem Set 1 – conversion of units.
       2. Problem Set 2 – conversion of grams to moles, moles to grams, atoms to moles, moles to atoms
       3. Problem Set 4 – determine the limiting reagent and the theoretical yield
       4. Problem Set 5 – determine the concentration of a solution and the concentration or a dilution
    c. Know and be able to apply the rules for significant figures.
       1. All non-zero digits are significant.
       2. Zeros between non-zero digits are significant.
       3. Zeros to the left of the first non-zero digit are not significant.
       4. If a number ends in zeros to the right of the decimal point, those zeros are significant.